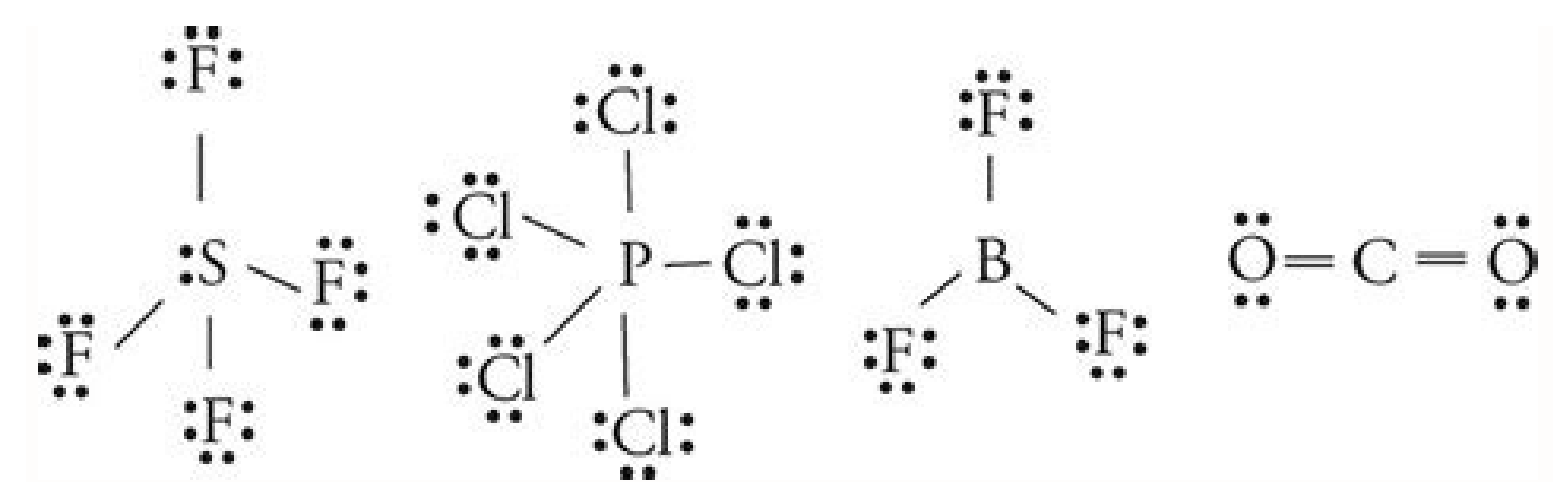
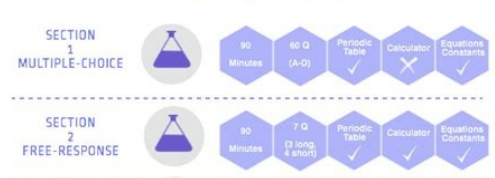


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AP CHEMISTRY 2014

Monday May 5th, 8.00 AM



OTHER NOTES ON EACH SECTION

MCQ - up to 50% in "lets" of 2-4 questions with a stem or stimulus
 FRQ - "long" approx. 15-20 mins each "short" approx. 7-10 mins each
 All FRQ's are multi-part

SCORING

MCQ	50
50% of final grade	No penalty for wrong answer
FRQ	50
50% of final grade	Questions weighted

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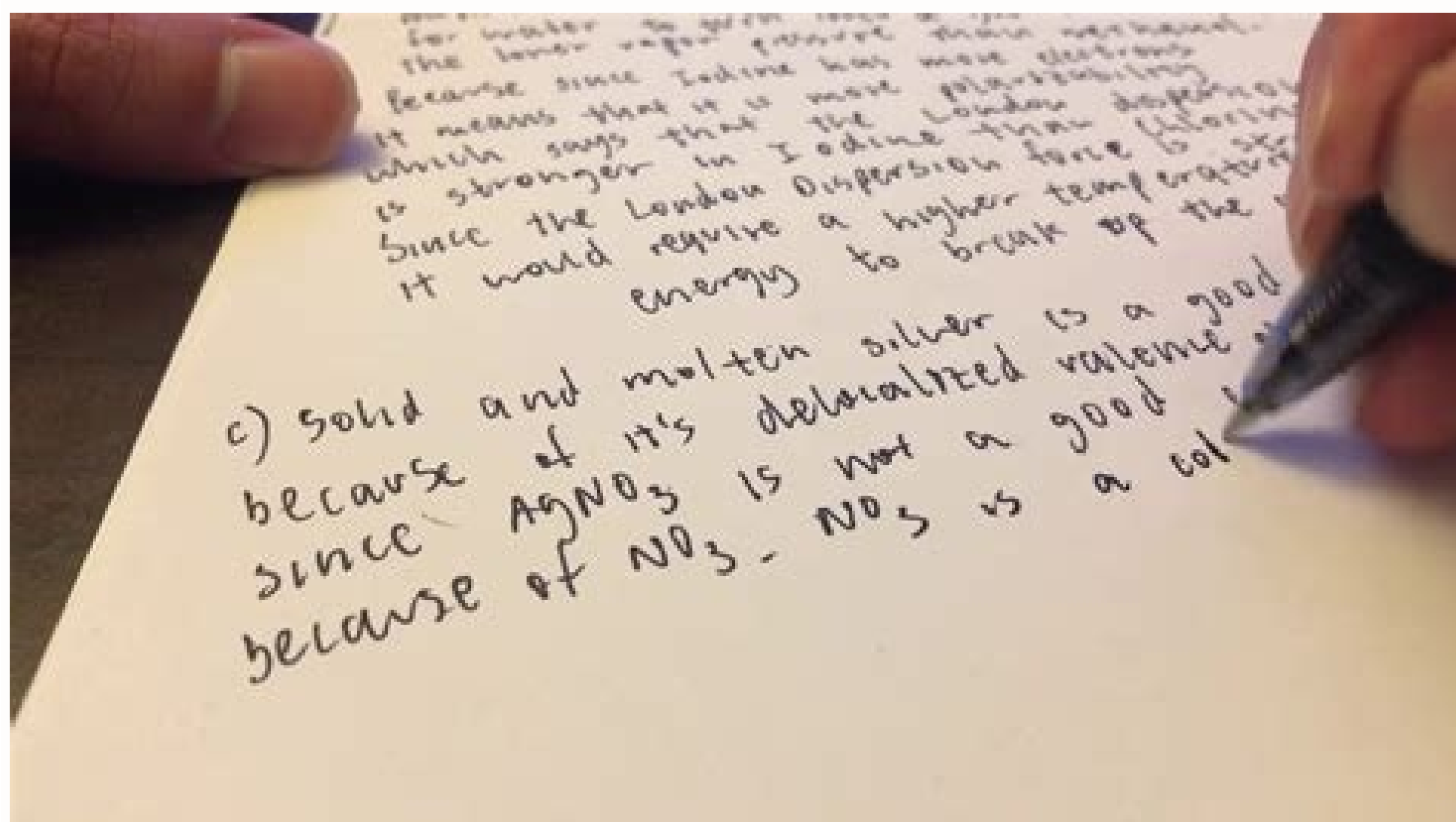




AP PREP FOR ADVANCED CHEMISTRY



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How to understand equilibrium chemistry. Ap chemistry unit 7 equilibrium test answers.

Learn about the fundamental concepts of chemistry including structure and states of matter, intermolecular forces, and reactions. You'll do hands-on lab investigations and use chemical calculations to solve problems. Note: Save your lab notebooks and reports; colleges may ask to see them before granting you credit. Designing experiments and procedures to test a prediction or theory Creating graphs, diagrams, and models that represent chemical phenomena Explaining how the microscopic structure of a substance determines its chemical properties Balancing a chemical equation Making a scientific claim and supporting it with evidence A one-year, introductory college general chemistry course High school courses in chemistry and Algebra II The course content outlined below is organized into commonly taught units of study that provide one possible sequence for the course. Your teacher may choose to organize the course content differently based on local priorities and preferences. Course Content Topics may include: Moles and molar mass Mass spectroscopy of elements Elemental composition of pure substances Composition of mixtures Atomic structure and electron configuration Photoelectron spectroscopy Periodic trends Valence electrons and ionic compounds Topics may include: Types of chemical bonds Intramolecular force and potential energy Structure of ionic solids Structure of metals and alloys Lewis diagrams Resonance and formal charge VSEPR and bond hybridization Topics may include: Intermolecular forces Solids, liquids, and gases Kinetic molecular theory Solutions and mixtures Photoelectric effect Topics may include: Introduction for reactions Net ionic equations Representations of reactions Physical and chemical changes Stoichiometry Types of chemical reactions Topics may include: Reaction rate Introduction to rate law Elementary reactions Collision model Introduction to reaction mechanisms Multistep reaction energy profile Catalysis Topics may include: Endothermic and exothermic processes Heat transfer and thermal equilibrium Heat capacity and calorimetry Energy of phase changes Introduction to enthalpy of reaction Enthalpy of formation Hess's law Topics may include: Introduction to equilibrium Calculating the equilibrium constant Calculating equilibrium concentrations Introduction to Le Châtelier's principle Introduction to solubility equilibria pH and solubility Free energy of dissolution Topics may include: Introduction to acids and bases pH and pOH of strong acids and bases Acid-base reactions and buffers Molecular structure of acids and bases pH and pKa Properties of buffers Topics may include: Introduction to entropy Gibbs free energy and thermodynamic favorability Thermodynamic and kinetic control Free energy and equilibrium Galvanic (voltaic) and electrolyte cells Electrolysis and Faraday's law The AP Chemistry Exam has consistent question types, weighting, and scoring guidelines every year, so you and your students know what to expect on exam day. Starting with the 2023 exam, a scientific or graphing calculator is recommended for use on both sections of the exam. Section I: Multiple Choice 60 Questions | 1 Hour 30 minutes | 50% of Exam Score Questions are either discrete questions or question sets, in which students are provided with a stimulus or a set of data and a series of related questions. Section II: Free Response 7 Questions | 1 Hour 45 Minutes | 50% of Exam Score There are 3 long-answer and 4 short-answer questions. Long-answer questions are worth 10 points each; short-answer questions are worth 4 points each. The questions assess all 6 of the course skills: models and representations, question and method, representing data and phenomena, model analysis, mathematical routines, and argumentation. Calculators are now allowed on both sections of the exam, starting with the 2023 exam, per the AP Exam Calculator Policy. Make sure you and your students know what calculators are permitted. View free-response questions and scoring information from past exams. Exam Overview The AP Chemistry Exam will test your understanding of the scientific concepts covered in the course units, as well as your ability to design and describe chemical experiments. Starting with the 2023 exam, a scientific or graphing calculator is recommended for use on both sections of the exam. You will be provided with a periodic table and a formula sheet that lists specific and relevant formulas for use on the exam. Exam Duration 3hrs 15mins Exam Components The multiple-choice section includes individual, single questions as well as sets of questions that refer to the same diagram or data presentation. Questions will test your ability to: Explain, analyze, and interpret models and representations of chemical properties or phenomena Design experiments and procedures to test a prediction or theory Solve problems using mathematical relationships Make or justify a scientific claim and support it with evidence and/or reasoning A scientific or graphing calculator is recommended for use on Section I The free-response section includes three long essay questions (worth 10 points each) and four short-answer questions (worth 4 points each). Questions will test your ability to: Explain, analyze, and interpret models and representations of chemical properties or phenomena Design experiments and procedures to test a prediction or theory Represent data and chemical phenomena with graphs and diagrams Solve problems using mathematical relationships Make or justify a scientific claim and support it with evidence and/or reasoning A scientific or graphing calculator is recommended for use on Section II. Students in AP Chemistry need to memorize a lot of information, but we have some good news! During the AP exam, you'll have access to a formula sheet that includes a periodic table as well as useful equations. However, you'll need to have a good sense of what's on the AP Chem formula sheet and how to use it in order to make the most of it during the test. In this guide, we explain everything you need to know about the AP Chemistry reference table. We go over what the formula sheet looks like, what it includes, what's important information it doesn't contain but that you should know, and how you can get the most out of it. What's on the AP Chem Formula Sheet? The AP Chemistry reference sheet is actually three pages of information. It will be at the beginning of both the multiple-choice and free-response sections of the test, so you'll have access to it for the entire AP Chemistry exam. This means you don't need to memorize any of the information the formula sheet contains. So what's on this AP Chem reference sheet? Check it out here on pages 2-4, as well as the screenshots below (click on each image to enlarge). The AP Chem equation sheet covers six main topics. Periodic table of elements Atomic structure Equilibrium Kinetics Gases, liquids, and solutions Thermodynamics/electrochemistry Periodic Table of Elements The AP Chem equation sheet includes a complete periodic table. For each element, it includes the element name, symbol, atomic number, and atomic mass. Atomic Structure 2 equations Values for: Planck's constant Speed of light Avogadro's number Electron charge Equilibrium Kinetics Gases, Liquids, and Solutions 9 equations Values for: Gas constant (3) 1 atm STP Ideal gas at STP Thermodynamics/Electrochemistry 9 equations Values for: What Isn't on the AP Chem Formula Sheet That You Need to Know for the Exam? While the AP Chemistry reference table does a good job of covering the majority of equations you need to know for the exam, there is some information it's missing. That's what we cover in this section. Be sure to know these equations for the exam. Rate Law One of the laws you learn in AP Chemistry is that, at a constant temperature, the rate of a chemical reaction depends only on the concentrations of the substances that influence the rate. These substances are typically reactants, but can also include products and catalysts. The rate law formula is: $\text{rate} = k[A]^m[B]^n$ [A] and [B] represent molar concentrations of reactants k is the rate constant (which varies depending on the reaction and the temperature) a and b are typically positive integers that must be found experimentally Rate law is used to estimate the relationship between the rate of a reaction and the concentrations of reactants. You may see it on questions related to kinetics. Coulomb's Law Coulomb's Law states that magnitude of electric force between the charges q1 and q2 is directly proportional to the magnitude of the charges and inversely proportional to the square of the distance between them. You may see it on electrochemistry questions. $F = k(q_1q_2/r^2)$ r = distance of separation Important Percentages You might need to use these equations on any portion of the test. They determine how accurate estimates are compared to actual results. $\% \text{error} = (\text{actual value} - \text{theoretical value}) / (\text{theoretical value}) * 100$ $\% \text{yield} = (\text{actual yield} / \text{theoretical yield}) * 100$ Looking for help studying for your AP exams? Our one-on-one online AP tutoring services can help you prepare for your AP exams. Get matched with a top tutor who got a high score on the exam you're studying for! 4 Tips for Getting the Most Out of the AP Chem Equation Sheet The AP Chemistry formula sheet can be extremely helpful during the exam, but only if you know how to use it properly. Here are four tips for getting the most out of it. #1: Know How to Use Every Equation on the AP Chem Formula Sheet The AP Chem reference sheet contains 32 equations, and you should know how each of them works before test day. The College Board isn't one for including extraneous information on formula sheets, so if a formula is included on the sheet, that means there's a solid chance you'll need to use it on the exam. You don't want to waste your exam trying to learn how to use, say Planck's equation, and hoping you got it right! Well before exam day, be sure to go through every formula on the sheet and make sure you understand it and know how to use it. The keys included in the formula sheet are also helpful for figuring out what different variables represent in the equations, so don't neglect them either. #2: Use Your Calculator Wisely You're allowed a graphing calculator on the free-response portion of the exam, but not the multiple-choice section. You can make good use of your calculator by programming constants and formulas into it so you can solve equations more easily. However, don't think this means you get out of understanding the actual formulas themselves. You still need to know when and how to use each equation on the AP Chemistry reference table, not to mention you won't have your calculator for multiple-choice questions. So know when your calculator is useful, but don't over-rely on it. #3: Take Practice Tests Using the Formula Sheet Answering practice questions and taking practice tests using the AP Chem formula sheet is key to doing well on the exam. You should take several AP Chemistry practice tests before the actual AP exam. For each of them, use the official formula sheet. Your teacher will also likely give you a copy of the AP Chem formula sheet for your in-class exams, so you can get some practice in there too. If you need help finding practice tests, check out our guide specifically on where to find the best AP Chemistry practice questions and tests. #4: Become an Expert on the Periodic Table If there's one guarantee for the AP Chemistry exam, it's that you'll need to look at the included periodic table during the test. The periodic table included in the AP Chemistry equation sheet has all the information you'll need to answer questions about it. However, that doesn't mean you can ignore it until it comes time to take the test. It's very important to be familiar with the periodic table throughout the year. You should know what each value in the table represents, how to use group numbers to determine the number of bonds and valence electrons an element can form, and what the key periodic trends are. Spend the most time on groups 1-8A, as you'll be tested on these elements most often. What's Next? Learn the difference between Physics 1, Physics 2, and Physics C in our guide to AP Physics classes. How many AP courses total should you take? What if you want to get into an Ivy League school? Read our advice on AP course load here. Want to improve your SAT score by 160 points or your ACT score by 4 points? We've written a guide for each test about the top 5 strategies you must be using to have a shot at improving your score. Download it for free now:

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